

A second triclinic polymorph of (1-ammonio-1-phosphonoethyl)-phosphonate

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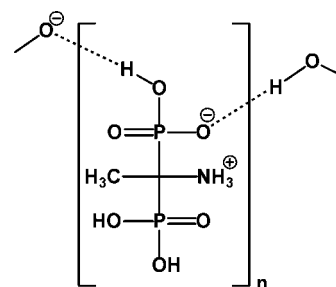
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.049; wR factor = 0.122; data-to-parameter ratio = 12.1.

The asymmetric unit of the second polymorph of the title compound, $\text{C}_2\text{H}_9\text{NO}_6\text{P}_2$, contains one molecule existing as a zwitterion. The N atom of the ammonio group is protonated and one of the phosphonic acid groups is deprotonated. Bond lengths and angles are similar in both polymorphs. Besides the differences in cell parameters, the most significant structural difference between this structure and that of the first polymorph [Dudko, Bon, Kozachkova, Tsarik & Pekhnyo (2008), *Ukr. Khim. Zh.* **74**, 104–106] is the presence of strong symmetric hydrogen bonds between neighbouring phosphonate groups. H atoms involved in these hydrogen bonds are located at inversion centres and $\text{O}\cdots\text{O}$ distances are observed in the range 2.458 (5)–2.523 (5) Å. These bonds and additional $\text{O}-\text{H}\cdots\text{O}$ and $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds interlink the molecules, giving a three-dimensional supramolecular network.

Related literature

For the original polymorph, see: Dudko *et al.* (2008). For similar bisphosphonates, see: Fernández *et al.* (2003); Li *et al.* (2009). For general background on the usage of organic diphosphonic acids as chelating agents in metal extraction and as drugs to prevent calcification and inhibit bone resorption, see: Matczak-Jon & Videnova-Adrabska (2005); Matkovskaya *et al.* (2001). For examples of symmetrical $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds, see Catti & Ferraris (1976); Meot-Ner (2005). For bond-length data, see: Allen *et al.* (1987).



Experimental

Crystal data

$\text{C}_2\text{H}_9\text{NO}_6\text{P}_2$

$M_r = 205.04$

Triclinic, $P\bar{1}$

$a = 5.5674$ (11) Å

$b = 5.9023$ (12) Å

$c = 11.385$ (2) Å

$\alpha = 82.334$ (10)°

$\beta = 82.145$ (9)°

$\gamma = 78.148$ (10)°

$V = 360.56$ (12) Å³

$Z = 2$

Mo $K\alpha$ radiation

$\mu = 0.59$ mm⁻¹

$T = 296$ K

$0.56 \times 0.16 \times 0.09$ mm

Data collection

Bruker APEXII CCD diffractometer

Absorption correction: multi-scan (*SADABS*; Bruker, 2007)

$T_{\min} = 0.734$, $T_{\max} = 0.949$

4331 measured reflections

1401 independent reflections

861 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.058$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.049$

$wR(F^2) = 0.122$

$S = 1.06$

1401 reflections

116 parameters

H atoms treated by a mixture of independent and constrained refinement

$\Delta\rho_{\text{max}} = 0.45$ e Å⁻³

$\Delta\rho_{\text{min}} = -0.43$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1}-\text{H1O}\cdots\text{O2}^{\text{i}}$	0.81 (6)	1.70 (6)	2.507 (5)	170 (6)
$\text{O3}-\text{H3O}\cdots\text{O4}^{\text{ii}}$	0.85 (7)	1.63 (7)	2.475 (5)	175 (7)
$\text{O5}-\text{H5O}\cdots\text{O5}^{\text{iii}}$	1.23 (1)	1.23 (1)	2.458 (5)	180 (0)
$\text{O6}-\text{H6O}\cdots\text{O6}^{\text{iv}}$	1.26 (1)	1.26 (1)	2.523 (6)	180 (1)
$\text{N1}-\text{H1A}\cdots\text{O2}^{\text{v}}$	0.90 (6)	2.11 (6)	2.929 (6)	150 (5)
$\text{N1}-\text{H1B}\cdots\text{O4}^{\text{vi}}$	0.86 (6)	2.06 (6)	2.896 (5)	165 (5)
$\text{N1}-\text{H1C}\cdots\text{O6}^{\text{ii}}$	0.82 (6)	2.50 (6)	3.196 (6)	145 (5)

Symmetry codes: (i) $-x+1, -y, -z$; (ii) $x, y-1, z$; (iii) $-x+1, -y, -z+1$; (iv) $-x, -y+1, -z+1$; (v) $x-1, y, z$.

Data collection: *APEX2* (Bruker, 2007); cell refinement: *S SAINT* (Bruker, 2007); data reduction: *S SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IM2293).

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supplementary materials

Acta Cryst. (2011). E67, o1651–o1652 [doi:10.1107/S1600536811022239]

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Comment

Diphosphonic acids with the P–C–P fragment, commonly named bisphosphonates, are well known since the 19th century. Due to the specific geometry and mutual influence of phosphonate fragments, compounds of this class possess a number of unique properties compared to other derivatives of phosphonic acids. Diphosphonic acids are structural analogues of inorganic pyrophosphate, one of the major metabolites in the cells, involved as a product in more than sixty biochemical reactions. Drugs prepared on the basis of bisphosphonates are highly efficient as a regulator of calcium metabolism and the immune response, they are used as anti-neoplastic, anti-inflammatory and antiviral agents and drugs with analgesic effect and as a component of toothpastes bisphosphonates prevent the formation of tartar (Matkovskaya *et al.*, 2001).

However, it is still not clearly understood why small structural modifications of the bisphosphonates may lead to extensive alterations in their physicochemical, biological and toxicological characteristics (Matczak-Jon & Videnova-Adrabinska, 2005). As a consequence of that determination of the structure of the bisphosphonates is very important to understand the influence of structural modifications on complex-forming abilities and physiological activities and deriving structure properties relations in general. In the present work we report a second polymorph of the title compound which crystallizes in the space group ($P\bar{1}$) whereas the previously described polymorph modification also crystallizes in the triclinic space group $P\bar{1}$ but with different cell parameters and cell volume (Dudko *et al.*, 2008). The asymmetric unit of the title compound (Fig. 1) contains one molecule in zwitterionic form with a proton transferred from one of the phosphonic groups to the amino group which is typical for all investigated 1-aminodiphosphonic acids (Fernández *et al.*, 2003; Li *et al.*, 2009). Bond lengths and angles are within normal ranges (Allen *et al.*, 1987) and are comparable with the first polymorph modification of the compound. It is generally accepted that, for O—H—O interactions where O···O is about 2.50 Å, examples can be found of truly symmetric hydrogen bonds, most of which have crystallographic equivalence between donor–acceptor atoms (Meot-Ner, 2005; Catti & Ferraris, 1976). The title compound, displays such short and strong symmetric hydrogen bonds between neighboring phosphonate groups, with the H atoms located at inversion centres and O···O distances of 2.458 (5) and 2.523 (5) Å. Multiple N—H···O and O—H···O hydrogen bonds in the crystal structure form an intricate three-dimensional supramolecular network (Fig.2, Table 1).

Experimental

The title compound C₂H₉NO₆P₂, was obtained by the reaction of acetonitrile which was brought into contact with the dry hydrogen chloride and phosphorus trichloride at 278 K followed by the dropwise addition of water. The obtained solution was treated by a mixture of acetone and diethyl ether, resulting in a white precipitate of the title compound. The resulting residue was dissolved in water and was stored in a dark place for slow evaporation. After 14 d of staying, suitable crystals for X-ray data collection were obtained.

Refinement

H atoms bonded to O and N atoms were located in a difference Fourier map. Their positions were refined freely whereas thermal parameters were fixed to $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{N},\text{O})$. Other H atoms which bonded to C were positioned geometrically and refined using a riding model with C—H = 0.96 Å for CH₃ with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$.

Figures

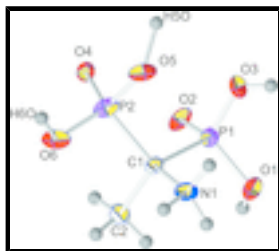


Fig. 1. View of the molecular structure of the title compound. Ellipsoids are drawn at the 50% probability level. H atoms are represented as small spheres of arbitrary radii.

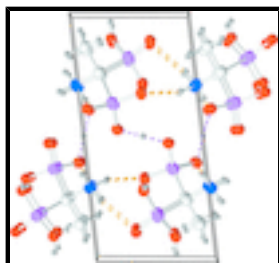


Fig. 2. Crystal packing of title compound, projection along *b* axis. Dashed pink lines indicate symmetrical hydrogen bonds while dashed yellow lines indicate ordinary hydrogen bonds.

(1-azaniumyl-1-phosphonoethyl)phosphonate

Crystal data

C₂H₉NO₆P₂

$M_r = 205.04$

Triclinic, *P*1

Hall symbol: -P 1

$a = 5.5674$ (11) Å

$b = 5.9023$ (12) Å

$c = 11.385$ (2) Å

$\alpha = 82.334$ (10)°

$\beta = 82.145$ (9)°

$\gamma = 78.148$ (10)°

$V = 360.56$ (12) Å³

$Z = 2$

$F(000) = 212$

$D_x = 1.889$ Mg m⁻³

Melting point: 551 K

Mo *K*α radiation, $\lambda = 0.71073$ Å

Cell parameters from 1315 reflections

$\theta = 3.6$ – 25.3 °

$\mu = 0.59$ mm⁻¹

$T = 296$ K

Needle, colourless

$0.56 \times 0.16 \times 0.09$ mm

Data collection

Bruker APEXII CCD
diffractometer

Radiation source: fine-focus sealed tube
graphite

1401 independent reflections

861 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.058$

φ and ω scans	$\theta_{\max} = 26.0^\circ$, $\theta_{\min} = 1.8^\circ$
Absorption correction: multi-scan (<i>SADABS</i> ; Bruker, 2007)	$h = -6 \rightarrow 6$
$T_{\min} = 0.734$, $T_{\max} = 0.949$	$k = -7 \rightarrow 7$
4331 measured reflections	$l = -12 \rightarrow 14$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.049$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.122$	H atoms treated by a mixture of independent and constrained refinement
$S = 1.06$	$w = 1/[\sigma^2(F_o^2) + (0.0387P)^2 + 0.6717P]$
1401 reflections	where $P = (F_o^2 + 2F_c^2)/3$
116 parameters	$(\Delta/\sigma)_{\max} < 0.001$
0 restraints	$\Delta\rho_{\max} = 0.45 \text{ e } \text{\AA}^{-3}$
	$\Delta\rho_{\min} = -0.43 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
P1	0.4512 (2)	-0.0726 (2)	0.18108 (11)	0.0243 (4)
P2	0.2798 (2)	0.2733 (2)	0.36756 (11)	0.0266 (4)
O1	0.3269 (6)	-0.1949 (6)	0.1024 (3)	0.0327 (9)
H1O	0.337 (11)	-0.133 (11)	0.034 (5)	0.049*
O2	0.6470 (5)	0.0492 (6)	0.1152 (3)	0.0337 (9)
O3	0.5438 (7)	-0.2366 (6)	0.2838 (4)	0.0475 (11)
H3O	0.531 (12)	-0.379 (12)	0.301 (6)	0.071*
O4	0.5260 (5)	0.3445 (5)	0.3268 (3)	0.0236 (8)
O5	0.2923 (6)	0.0865 (7)	0.4724 (3)	0.0344 (9)
H5O	0.5000	0.0000	0.5000	0.052*
O6	0.0723 (6)	0.4792 (7)	0.3914 (3)	0.0421 (10)
H6O	0.0000	0.5000	0.5000	0.063*

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N1	-0.0104 (8)	0.0192 (9)	0.2928 (4)	0.0287 (11)
H1A	-0.086 (10)	-0.025 (9)	0.237 (5)	0.043*
H1B	-0.151 (10)	0.102 (10)	0.316 (5)	0.043*
H1C	0.030 (10)	-0.091 (10)	0.342 (5)	0.043*
C1	0.1979 (8)	0.1433 (8)	0.2439 (4)	0.0195 (10)
C2	0.1069 (9)	0.3297 (8)	0.1441 (4)	0.0278 (12)
H2A	0.0827	0.2556	0.0776	0.042*
H2B	0.2276	0.4265	0.1191	0.042*
H2C	-0.0464	0.4235	0.1730	0.042*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
P1	0.0236 (7)	0.0195 (7)	0.0316 (7)	-0.0033 (5)	-0.0072 (6)	-0.0066 (6)
P2	0.0204 (7)	0.0387 (9)	0.0230 (7)	-0.0080 (6)	-0.0030 (5)	-0.0074 (6)
O1	0.038 (2)	0.028 (2)	0.037 (2)	-0.0124 (17)	-0.0077 (18)	-0.0095 (17)
O2	0.0203 (18)	0.045 (2)	0.042 (2)	-0.0155 (16)	0.0069 (16)	-0.0208 (18)
O3	0.065 (3)	0.016 (2)	0.068 (3)	-0.0036 (19)	-0.042 (2)	0.002 (2)
O4	0.0182 (16)	0.0200 (18)	0.0344 (19)	-0.0022 (13)	-0.0056 (14)	-0.0091 (15)
O5	0.0272 (19)	0.055 (2)	0.0241 (18)	-0.0172 (17)	-0.0077 (15)	0.0035 (17)
O6	0.0231 (19)	0.067 (3)	0.034 (2)	0.0124 (18)	-0.0057 (16)	-0.0270 (19)
N1	0.017 (2)	0.036 (3)	0.033 (3)	-0.009 (2)	-0.006 (2)	0.005 (2)
C1	0.016 (2)	0.022 (3)	0.022 (2)	-0.0064 (19)	-0.0040 (19)	0.000 (2)
C2	0.031 (3)	0.023 (3)	0.027 (3)	0.000 (2)	-0.007 (2)	-0.001 (2)

Geometric parameters (\AA , $^\circ$)

P1—O2	1.490 (3)	O5—H5O	1.229 (0)
P1—O3	1.493 (4)	O6—H6O	1.261 (0)
P1—O1	1.533 (4)	N1—C1	1.502 (6)
P1—C1	1.831 (5)	N1—H1A	0.90 (6)
P2—O4	1.510 (3)	N1—H1B	0.86 (6)
P2—O5	1.512 (4)	N1—H1C	0.82 (6)
P2—O6	1.520 (3)	C1—C2	1.534 (6)
P2—C1	1.840 (4)	C2—H2A	0.9600
O1—H1O	0.81 (6)	C2—H2B	0.9600
O3—H3O	0.85 (7)	C2—H2C	0.9600
O2—P1—O3	111.7 (2)	C1—N1—H1B	118 (4)
O2—P1—O1	114.3 (2)	H1A—N1—H1B	87 (4)
O3—P1—O1	110.8 (2)	C1—N1—H1C	113 (4)
O2—P1—C1	109.0 (2)	H1A—N1—H1C	110 (5)
O3—P1—C1	106.5 (2)	H1B—N1—H1C	111 (6)
O1—P1—C1	103.86 (19)	N1—C1—C2	107.8 (4)
O4—P2—O5	112.46 (18)	N1—C1—P1	107.2 (3)
O4—P2—O6	112.8 (2)	C2—C1—P1	109.4 (3)
O5—P2—O6	111.7 (2)	N1—C1—P2	107.4 (3)
O4—P2—C1	107.07 (18)	C2—C1—P2	111.7 (3)
O5—P2—C1	106.2 (2)	P1—C1—P2	113.1 (2)

O6—P2—C1	105.98 (19)	C1—C2—H2A	109.5
P1—O1—H1O	110 (4)	C1—C2—H2B	109.5
P1—O3—H3O	128 (4)	H2A—C2—H2B	109.5
P2—O5—H5O	115.9 (0)	C1—C2—H2C	109.5
P2—O6—H6O	115.1 (0)	H2A—C2—H2C	109.5
C1—N1—H1A	115 (3)	H2B—C2—H2C	109.5
O2—P1—C1—N1	171.5 (3)	O4—P2—C1—N1	166.1 (3)
O3—P1—C1—N1	-67.9 (3)	O5—P2—C1—N1	45.8 (3)
O1—P1—C1—N1	49.2 (3)	O6—P2—C1—N1	-73.2 (4)
O2—P1—C1—C2	54.9 (4)	O4—P2—C1—C2	-76.0 (3)
O3—P1—C1—C2	175.5 (3)	O5—P2—C1—C2	163.7 (3)
O1—P1—C1—C2	-67.4 (3)	O6—P2—C1—C2	44.7 (4)
O2—P1—C1—P2	-70.3 (3)	O4—P2—C1—P1	48.0 (3)
O3—P1—C1—P2	50.3 (3)	O5—P2—C1—P1	-72.4 (3)
O1—P1—C1—P2	167.4 (2)	O6—P2—C1—P1	168.7 (2)

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1O \cdots O2 ⁱ	0.81 (6)	1.70 (6)	2.507 (5)	170 (6)
O3—H3O \cdots O4 ⁱⁱ	0.85 (7)	1.63 (7)	2.475 (5)	175 (7)
O5—H5O \cdots O5 ⁱⁱⁱ	1.23 (1)	1.23 (1)	2.458 (5)	180 (0)
O6—H6O \cdots O6 ^{iv}	1.26 (1)	1.26 (1)	2.523 (6)	180.(1)
N1—H1A \cdots O2 ^v	0.90 (6)	2.11 (6)	2.929 (6)	150 (5)
N1—H1B \cdots O4 ^v	0.86 (6)	2.06 (6)	2.896 (5)	165 (5)
N1—H1C \cdots O6 ⁱⁱ	0.82 (6)	2.50 (6)	3.196 (6)	145 (5)

Symmetry codes: (i) $-x+1, -y, -z$; (ii) $x, y-1, z$; (iii) $-x+1, -y, -z+1$; (iv) $-x, -y+1, -z+1$; (v) $x-1, y, z$.

Fig. 1

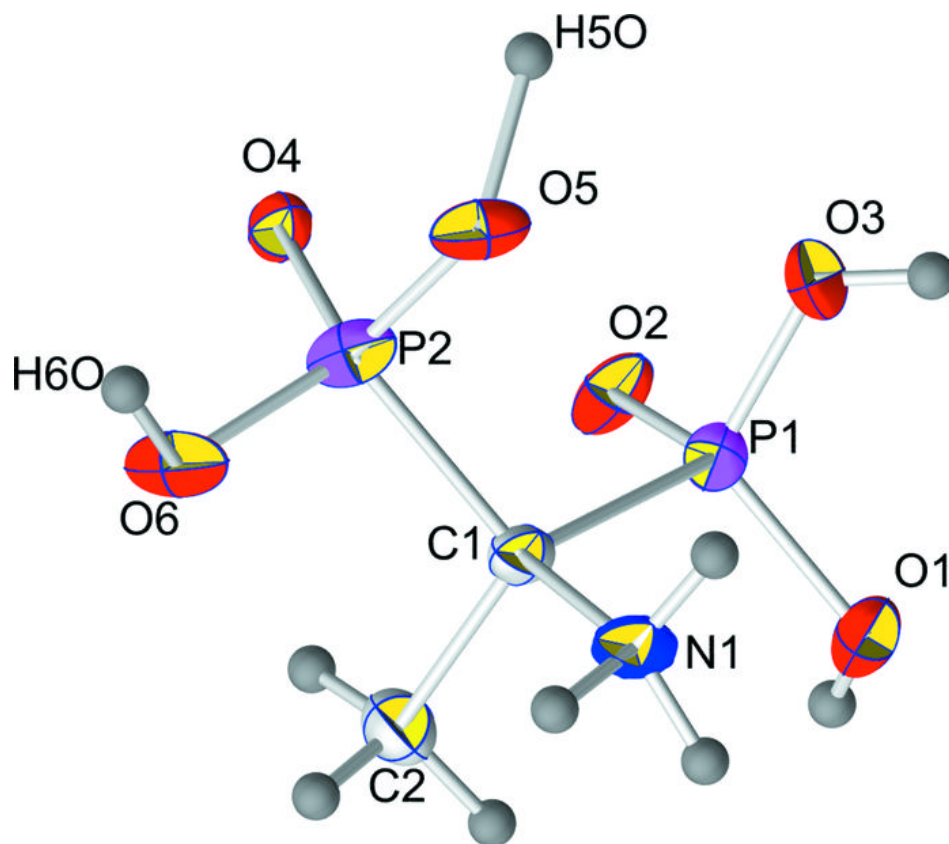


Fig. 2

